

DIVISION XII/COMMISSION 14/WORKING GROUP ON MOLECULAR DATA

CHAIR	Steven R. Federman
VICE-CHAIR	Peter F. Bernath
VICE-CHAIR	Holger S.P. Müller

TRIENNIAL REPORT 2009-2012

1. Introduction

The current report covers the period from the second half of 2003 to the first half of 2011, bringing the Working Group's efforts up to date, and is divided into three main sections covering rotational, vibrational, and electronic spectroscopy. Rather than being exhaustive, space limitations only allow us to highlight a representative sample of work on molecular spectra. Related research on collisions, reactions on grain surfaces, and astrochemistry appear in the report by another Working Group. These also recount recent conferences and workshops on molecular astrophysics.

2. Rotational Spectra

A large number of reviews have appeared dealing with rotational spectra of molecules potentially relevant to radio-astronomical observations. Therefore, emphasis is placed on investigations dealing with molecular species already observed in space. Related molecules are included to a large extent also. The following groups have been created – hydride species, anions, molecules which may occur in circumstellar envelopes of late type stars, complex molecules, weed species, and other molecules – and are discussed in turn.

Several databases provide rotational spectra of molecular species of astrophysical and astrochemical relevance. The two most important sources for predictions generated from experimental data by employing appropriate Hamiltonian models are the Cologne Database for Molecular Spectroscopy, CDMS (<http://www.astro.uni-koeln.de/cdms/>) with its catalog (<http://www.astro.uni-koeln.de/cdms/catalog>). An updated description appeared in Müller et al. (2005) and the JPL catalog (<http://spec.jpl.nasa.gov/>). Both also provide primary information, i.e. laboratory data with uncertainties, mostly in special archive sections. Additional primary data are available in the Toyama Microwave Atlas (<http://www.sci.u-toyama.ac.jp/phys/4ken/atlas/>). A useful resource on the detection of certain molecular transitions in space is the NIST Recommended Rest Frequencies for Observed Interstellar Molecular Microwave Transitions, which has been updated and described by Lovas (2004).

The European FP7 project Virtual Atomic and Molecular Data Centre, VAMDC, (<http://www.vamdc.org/>) aims at combining several spectroscopic, collisional, and kinetic databases. The CDMS is the rotational spectroscopy database taking part; several infrared databases are also involved. The project has been described by Dubernet et al. (2010). Other tertiary sources combining data from various databases are, e.g., Cassis

(<http://cassis.cesr.fr/>), which provides tools to analyze astronomical spectra, or splatalogue (<http://www.splatalogue.net/>).

2.1. Hydrides

Hydrides here are all molecules consisting of one non-metal atom and one or more H atoms. They may be neutral or positively charged. Metal hydrides will be dealt with in subsection 2.3. Rotational spectra of some anionic hydride molecules have been obtained also, but such molecules have not yet been detected in space, see also subsection 2.2. Hydrides are usually difficult to observe from the ground as their fundamental transitions occur in the upper millimeter region for heavier species to the terahertz region for lighter ones. Some hydrides have recently been detected from the ground, such as $^{13}\text{CH}^+$ with the CSO, or SH^+ and OH^+ with the APEX telescope. The launch of the *Herschel* satellite in May 2009 has created a wealth of opportunities to investigate hydrides, in particular with the high resolution instrument HIFI. H_2O^+ , H_2Cl^+ , ND, and possibly HCl^+ have been detected, and the fundamental transitions of CH^+ and of HF have been observed for the first time.

The very light hydrides H_2D^+ and HD_2^+ have attracted considerable attention (Amano & Hirao 2005; Asvany et al. 2008; Yonezu et al. 2009). Particularly noteworthy is the measurement of the $1_{0,1} - 0_{0,0}$ transition of H_2D^+ more than 60 MHz away from the previously accepted value by Asvany et al. (2008). A similarly large deviation occurred between the initial transition frequency of the $J = 1 - 0$ transition of CH^+ (Pearson & Drouin 2006) and the one measured by Amano (2010a), who also recorded the same transition for $^{13}\text{CH}^+$ and CD^+ . Müller (2010) used the latter transition frequencies together with data from electronic spectra to derive extensive predictions for rotational and rovibrational transitions of several CH^+ isotopologues. Isotopic CH (Halfen et al. 2008), CH_2 (Brünken et al. 2005), CHD (Ozeki et al. 2011), CH_2D^+ (Amano 2010b), and CH_3D (Drouin et al. 2009) have been other carbon containing hydrides studied.

No accurate rotational transition frequencies are available for CH_2^+ or NH_2^+ . Their fundamental transitions are beyond the region which can be studied with HIFI.

The nitrogen hydrides NH (Flores-Mijangos et al. 2004), NH^+ (Hübers et al. 2009), NH_3 (Yu et al. 2010), as well as the isotopologues NHD_2 (Endres et al. 2006), $^{15}\text{NH}_2\text{D}$ and $^{15}\text{NHD}_2$ (Elkeurti et al. 2008) have been investigated recently. Transition frequencies have been determined for the oxygen hydrides ^{17}OH (Polehampton et al. 2003), OH^+ (Müller et al. 2005), H_2O and H_2^{18}O (Golubiatnikov et al. 2006), H_2^{17}O (Puzzarini et al. 2009), D_2O (Brünken et al. 2007a; Cazzoli et al. 2010a), H_3O^+ (Yu et al. 2009), and H_2DO^+ (Müller et al. 2010).

Halogen containing molecules have attracted increased attention in recent years. Among the hydride species investigated recently are DF (Cazzoli et al. 2006a), HF^+ (Allen et al. 2004), H_2F^+ (Fujimori et al. 2011), and H_2Cl^+ (Araki et al. 2001).

Other investigations of heavier hydrides include Lamb-dip measurements of low-lying rotational states of PH_3 (Cazzoli & Puzzarini 2006) and a combined analysis of spectroscopic data of SH^+ (Brown & Müller 2009).

2.2. Anions

There had been speculation about the possibility of molecular anions in space for quite a while. However, accurate transition frequencies were only known for OH^- and SH^- ; for a few other species approximate values were known from infrared spectroscopy. However, all these species were deemed to be too fragile with respect to photolysis. Moreover, all

these species were hydride species with fundamental transitions not easily accessible from the ground.

Interestingly, a molecular line survey of the circumstellar envelope of the carbon-rich star CW Leo, also known as IRC +10216, between 28 and 50 GHz revealed several series of unidentified lines (Kawaguchi et al. 1995). One of these turned out to be caused by the anion C_6H^- (McCarthy et al. 2006). This finding sparked a flurry of investigations into laboratory rotational spectra of related molecular anions and the search for these species in space. C_2H^- (Brünken et al. 2007b), as well as C_4H^- and C_8H^- (Gupta et al. 2007), were characterized and, except for the smallest member, found in space. The isoelectronic molecules CN^- (Gottlieb et al. 2007) and C_3N^- (Thaddeus et al. 2008) were found in the lab as well as in the circumstellar envelope of CW Leo. The identification of C_5N^- also in that source seemed to be certain enough even in the absence of laboratory spectral data. Furthermore, submillimeter transitions have been recorded for CN^- , C_2H^- , and C_4H^- (Amano 2008) as well as for C_3N^- (Amano 2010c). In addition, the laboratory rotational spectrum of NCO^- was recently reported (Lattanzi et al. 2010).

2.3. Circumstellar molecules

The chemistry of circumstellar envelopes of late-type stars is special as a number of molecules have only been detected or are particularly abundant in such sources. Probably the most fascinating object in this regard is CW Leo. Whereas many astronomers have thought this source has a rather special chemistry, it turned out that it is a rather common source. The reason why molecules are particularly easily detected in this source is its proximity. For instance, several metal containing molecules have been detected first in the envelope of CW Leo, but in recent years MgNC, NaCN, and NaCl have been detected also in the envelopes of other AGB stars, the latter even in those of O-rich AGB stars such as IK Tau and VY CMa. Interestingly, recently the metal-containing molecules AlO and AlOH as well as PO have been detected first in the envelope of VY CMa. In addition, KCN and FeCN were detected toward CW Leo.

For a number of metal-containing molecules rotational spectra have been studied. Hydride species include AlH (Halfen & Ziurys 2004), CrH (Harrison et al. 2006), MnH (Halfen & Ziurys 2008), and FeH (Brown et al. 2006). Among the nitride and oxide species studied are VN and VO (Flory & Ziurys 2008), CoO (McLamarrah et al. 2005), ZnO (Zack et al. 2009), and even TiO_2 (Kania et al. 2011). Several halogenides have been investigated, such as NaCl and KCl (Caris et al. 2004), TiF (Sheridan et al. 2003), TiCl (Maeda et al. 2001), VCl (Halfen et al. 2009b), CoF (Harrison et al. 2007), CoCl (Flory et al. 2004), ZnF (Flory et al. 2006), and ZnCl (Tenenbaum et al. 2007). Also recorded were rotational spectra of CoCN and NiCN (Sheridan et al. 2004; Sheridan & Ziurys 2003).

The study of the rotational spectrum of C_2P (Halfen et al. 2009a) laid the ground work for its detection in the circumstellar envelope of CW Leo. This detection, in turn, suggests that molecules such as C_2F , C_2Cl , C_3F , and C_3Cl may be detectable as well. Rotational spectra have been detected for all of these species except for the first one (Sumiyoshi et al. 2003; Yoshikawa et al. 2009b,a).

Other molecules investigated, including predominantly or potentially circumstellar ones, are C_2H and C_6H in excited vibrational states (Killian et al. 2007; Gottlieb et al. 2010) and CH_3CP (Bizzocchi et al. 2003).

2.4. *Complex molecules*

Some of the smaller complex molecules with many emission or absorption lines observable by radio-astronomical means are presented in subsection 2.5.

Among the larger saturated or almost saturated molecules, propenal, propanal, and acetamide have been detected with the GBT in the microwave region. Interestingly, propylene has been detected with the IRAM 30 m telescope toward TMC-1. Three complex molecules have been detected with the same instrument in the course of a molecular line survey of Sagittarius B2(N) at 3 mm. These are aminoacetonitrile, *n*-propyl cyanide, and ethyl formate. The detection articles also feature critical evaluations of the spectroscopic parameters of the former two molecules (Belloche et al. 2008, 2009). The rotational spectrum of ethyl formate has also been revisited (Medvedev et al. 2009). Other studies include *cyclo*-propyl cyanide (Bizzocchi et al. 2008) and *iso*-propyl cyanide (Müller et al. 2011), two conformers of ethylene glycol (Christen & Müller 2003; Müller & Christen 2004), microwave spectra of several conformers of 1,2- and 1,3-propanediol (Lovas et al. 2009; Plusquellic et al. 2009), propane (Drouin et al. 2006), millimeter wave spectra of the amino acids glycine and alanine (Ilyushin et al. 2005; Hirata et al. 2008), acetamide (Ilyushin et al. 2004), methylamine (Ilyushin & Lovas 2007), acetic acid (Ilyushin et al. 2008), *n*-propanol (Kisiel et al. 2010), and diethyl ether (Walters et al. 2009). Other investigations may turn out to also be relevant once telescope arrays such as EVLA, NOEMA, or ALMA conduct large scale line surveys or dedicated searches for particular molecules.

2.5. *Weed species*

The term “weed species” has been coined for molecules that have very many emission or absorption lines in various sources, but considered mostly for star-forming regions. It should be emphasized that the plethora of lines is not only a nuisance, but there is also considerable information about, for example, temperature or density in these lines. The molecule with particularly many rather strong lines is methanol. It has been extensively studied up to the second torsional state (Xu et al. 2008); even higher states have been studied to some extent (Pearson et al. 2009), but these are difficult to model at present. CH₃OH has also been proposed as a particularly well suited molecule to investigate the possibility of temporal or spatial variations of fundamental constants (Jansen et al. 2011; Levshakov et al. 2011). One requirement is the knowledge of very highly accurate transition frequencies. Some of these have been summarized as well as newly reported ones by Müller et al. (2004). Other methanol isotopologues studied recently include CH₂DOH (Mukhopadhyay et al. 2002; Lauvergnat et al. 2009), CH₃OD (Duan et al. 2003), and CH₃¹⁸OH (Fisher et al. 2007).

Other weed molecules, for which several isotopic species have been investigated, are methyl cyanide (Müller et al. 2009), ethyl cyanide (Brauer et al. 2009; Demyk et al. 2007; Margulès et al. 2009b), vinyl cyanide (Kisiel et al. 2009), methyl formate (Ilyushin et al. 2009; Carvajal et al. 2009, 2010; Margulès et al. 2009a, 2010), formamide (Kryvda et al. 2009), and acetone (Groner et al. 2008; Lovas & Groner 2006). Moreover, dimethyl ether (Endres et al. 2009), its isomer ethanol (Pearson et al. 2008), and the only inorganic weed species SO₂ (Müller & Brünken 2005) have been studied.

One should keep in mind that weed species are not only seen in the ground vibrational state, but often in several excited states, and this is true for minor isotopic species as well. While deriving an appropriate Hamiltonian model is usually the most compact form to represent the rotational spectrum of a molecule in a given vibrational state, this task can be formidable and worse in cases of strong vibration-rotation interaction that

may take much more than months to be tackled. As an alternative, the De Lucia group proposed recording spectra at different temperatures. While this procedure yields huge amounts of data, and extrapolation in frequency or to higher temperatures is impossible and extrapolation to lower temperatures only to some extent, it seems to be still a pragmatic approach for some molecules. Several reports have been published on ethyl cyanide (Fortman et al. 2010) and one on vinyl cyanide (Fortman et al. 2011).

2.6. Other molecules

Other molecules, for which rotational spectra have been (re-) investigated, include cations such as HCNH^+ and CH_3CNH^+ (Amano et al. 2006), CS^+ (Bailleux et al. 2008), CF^+ (Cazzoli et al. 2010b), ^{15}N isotopologues of N_2H^+ (Dore et al. 2009), HCS^+ (Margulès et al. 2003), and HCO^+ (Tinti et al. 2007).

Other short-lived molecules include DNC (Brünken et al. 2006), C_3H and ^{13}C isotopologues of C_nH with $n = 3, 5 - 7$ (Caris et al. 2009; McCarthy & Thaddeus 2005), HOCN, HONC, and HSCN (Brünken et al. 2009a; Mladenović et al. 2009; Brünken et al. 2009b) as well as SiN and PN (Bizzocchi et al. 2006; Cazzoli et al. 2006b).

In addition, several stable or fairly stable molecules have been studied, including isotopologues of CO in their ground vibrational states (Puzzarini et al. 2003) and the main isotopologue in excited vibrational states (Gendriesch et al. 2009), various isotopologues in the ground and excited states of CS and SiS (Müller et al. 2005, 2007), isotopologues of HCN in their ground vibrational states (Cazzoli & Puzzarini 2005) and the main isotopologue in excited vibrational states (Zelinger et al. 2003), H_2CO (Brünken et al. 2003), the isoelectronic H_2CNH (Dore et al. 2010), H_2CS (Maeda et al. 2008), isotopologues of HCOOH (Lattanzi et al. 2008), and $\text{CH}_3\text{C}_2\text{H}$ and $\text{CH}_3\text{C}_4\text{H}$ (Cazzoli & Puzzarini 2008).

3. Vibrational Spectra

Here we describe vibration-rotation spectra of gaseous molecules of astronomical or potential astronomical interest. It is based in part on a review article entitled “Molecular astronomy of cool stars and sub-stellar objects” aimed at physical chemists and laboratory spectroscopists (Bernath 2009).

In addition to the references to particular molecules given below there are a number of spectral database compilations that are useful. Perhaps the most helpful is the HITRAN database that contains vibration-rotation line parameters for a large number of species such as H_2O , CO_2 , CO, etc. found in the Earth’s atmosphere (Rothman et al. 2009). HITRAN is widely used for astronomical applications although it is not always suitable because of missing lines and bands, particularly in the near infrared region. For example, “cool” astronomical objects such as brown dwarfs can have surface temperatures in excess of 1000 K, and HITRAN is designed for temperatures near 300 K. In this regard, there is a HITEMP database (Rothman et al. 2010) for H_2O , CO_2 , CO, NO, and OH that is more suitable for high temperature sources such as stellar atmospheres.

For larger molecules, individual vibration-rotation lines are no longer clearly resolved and it becomes necessary to replace line-by-line calculations by absorption cross sections. The main drawback to using cross sections is that a considerable number of laboratory measurements are needed to match the temperature and pressure conditions of the objects under observation. HITRAN also includes a number of high resolution infrared absorption cross sections for organic molecules such as ethane and acetone, but the broadening gas in HITRAN is air rather than H_2 , N_2 , or CO_2 . While the GEISA

database has significant overlap with HITRAN, it contains additional molecules of interest for studies of planetary atmospheres (Jacquinet-Husson et al. 2011). A very useful set of infrared absorption cross sections for several hundred molecules have been measured at the Pacific Northwest National Laboratory (PNNL) for the 600-6500 cm^{-1} (1.54–16.7 μm) range (Sharpe et al. 2004). The PNNL IR database, however, may not be completely suitable for astronomical applications, for example for planetary atmospheres. All PNNL spectra are recorded at relatively low resolution (0.112 cm^{-1}) as mixtures with pure nitrogen gas at pressures of 760 Torr and temperatures of 278, 293, or 323 K. Nevertheless, in the absence of spectra recorded under more appropriate experimental conditions, they can be very useful.

Other interesting general sources for infrared data are various high resolution spectral atlases of the Sun (Livingston & Wallace 2003; Hase et al. 2010) and sunspots (Wallace et al. 2001, 2002) because they include molecular (and atomic) line assignments. The web site spectrafactory (Cami et al. 2010a) is also useful for calculating infrared spectra of astronomical interest, although not all of the input line parameters are the most recent or recommended ones.

3.1. Diatomic Molecules

Diatomic molecules, particularly diatomic hydrides of the more abundant elements, are often observed through their infrared spectra. Molecular hydrogen itself is difficult to observe in the infrared because all transitions are electric-dipole forbidden; however, HD has a small dipole moment and very recently the 2–0 band was measured near 1.4 μm by cavity ring down spectroscopy (Kassi & Campargue 2011). In the case of OH, NH, and CH, the new infrared solar atlas (Hase et al. 2010) measured by the Atmospheric Chemistry Experiment (ACE) infrared Fourier transform spectrometer from orbit (Bernath et al. 2005) has led to an improvement in the spectroscopy. The infrared Fraunhofer lines were combined with laboratory measurements to extend the measured line positions to higher vibrational and rotational quantum numbers for OH (Bernath & Colin 2009), NH (Ram & Bernath 2010), and CH (Colin & Bernath 2010).

Laboratory vibration-rotation spectra for a considerable number of metal hydrides are known with recent measurements of, for example, BeH (Shayesteh et al. 2003a), MgH (Shayesteh et al. 2004a), and CaH (Shayesteh et al. 2004b); even the metal dihydrides BeH₂ (Shayesteh et al. 2003b) and MgH₂ (Shayesteh et al. 2003c) have been detected in the laboratory. Although metal hydrides such as MgH and CaH are detected in stellar atmospheres by their electronic transitions, there have been no infrared astronomical observations. Other diatomics such as CO, SO, SiO, HF, HCl, SH, CS, and SiS are commonly seen by infrared observations of cool stellar atmospheres (Bernath 2009); their spectroscopy however has not been improved much in recent years. An exception to this is some work on the line intensities of SiS (Cami et al. 2009) and HCl (Li et al. 2011).

3.2. Small Polyatomics

The spectra of “cold” water, ammonia, and methane as given in the HITRAN database are generally satisfactory for astronomical purposes, except for overtone and combination bands of NH₃ and CH₄ in the near infrared and visible regions. For spectra of hot samples as in brown dwarfs and exoplanets, the situation is much less sanguine, except in the case of water for which there has been extensive experimental (Zobov et al. 2008) and theoretical work (Barber et al. 2006). Much work is continuing on highly-excited water levels near dissociation and on computing line intensities ab initio, but the BT2 line list (Barber et al. 2006) is generally suitable for computing water opacities. For hot ammonia

there has been rapid recent progress with new laboratory spectra recorded (Hargreaves et al. 2011) and with at least two groups providing rather good calculated spectra (Huang et al. 2011; Yurchenko et al. 2011). Ammonia is seen in brown dwarfs and is thought to be the key molecule in defining a new class of ultracool Y-type dwarfs (Cushing et al. 2011). Methane is the laggard because of the difficulty of the problem, and the existing experimental data (Nassar & Bernath 2003) and calculations are not very satisfactory for simulating brown dwarf and exoplanet spectra. For CO₂, HITRAN for cold molecules and HITEMP or CDSD-4000 databases (Tashkun & Perevalov 2011) for hot molecules are recommended.

3.3. Large Molecules

The discovery of C₆₀ and C₇₀ in the young planetary nebula Tc1 was reported by Cami et al. (2010b) using the Spitzer Space Telescope. The discovery was rapidly confirmed with detections in other planetary nebulae; some of these sources also have spectral features generally attributed to polycyclic aromatic hydrocarbons (PAHs). Infrared emission spectra of gaseous C₆₀ and C₇₀ were recorded as a function of temperature by Nemes et al. (1994) and more recently the integrated molar absorptivity of the infrared bands of the solid has been measured as a function of temperature (Iglesias-Groth et al. 2011). There have also been extensive laboratory measurements and quantum chemical calculations of PAH spectra, some of which is summarized in the NASA Ames database (Bauschlicher et al. 2010). Traditionally infrared spectra of neutral and ionized PAHs have been recorded by matrix isolation spectroscopy, but now gas-phase measurements are possible (Ricks et al. 2009; Galu e et al. 2011). Computational studies of PAHs and related species have also advanced, and the quality and quantity of calculations available is remarkable (Hudgins et al. 2005; Bauschlicher et al. 2010).

4. Electronic Spectra

In this section we describe recent work on electronic spectra that includes line identification, energy levels, as well as data needed for photochemical models. The data come in a variety of forms, such as absorption cross sections (or equivalently oscillator strengths), predissociation widths, and analyses of line anomalies resulting from perturbations between energy levels. Here we can only present a representative sampling of work on electronic spectra. The number of experiments on electron excitation/scattering is great, and we provide a few illustrative examples of research in this area. This section is divided into three topics: interstellar matter, which includes diffuse molecular clouds and disks around newly formed stars as well as comets whose chemistry is similar, metal hydrides and oxides in the spectra of late-type stars, and the atmospheres of planets and their satellites. Recent attempts to identify the diffuse interstellar bands is not included, but instead we refer the reader to a review on laboratory astrophysics (Savin et al. 2011) that includes a discussion of this topic.

4.1. Interstellar matter

There is continued interest in spectroscopic studies of CO and much of the recent efforts are needed for improved photochemical modeling. A comprehensive analysis of Rydberg states has been published (Eidelsberg et al. 2004a), as has a compilation of triplet and singlet transitions (Eidelsberg & Rostas 2003) that included oscillator strengths. Improved wavelengths have appeared for the *A* – *X* system of bands in C¹⁷O and C¹⁸O (Steinmann et al. 2003; du Plessis et al. 2006, 2007), for the *E* – *X* (0,0) band in CO, ¹³CO, and ¹³C¹⁸O (Cacciani & Ubachs 2004), and for triplet-singlet bands involving the *e*, *d*, and *a*'

states in several isotopologues (du Plessis et al. 2007; Yang et al. 2008; Dickenson et al. 2010). Several quantum mechanical calculations of potential energy curves (Chakrabarti & Tennyson 2006; Vázquez et al. 2009; Lefebvre-Brion et al. 2010) have appeared. Empirical determinations of oscillator strengths for Rydberg transitions have been published, including one based on interstellar spectra acquired with the *Far Ultraviolet Spectroscopic Explorer* (Sheffer et al. 2003), two using synchrotron radiation (Eidelsberg et al. 2004b; Eidelsberg et al. 2006), and one employing electron scattering techniques (Kawahara et al. 2008). Gilijamse et al. (2007) have obtained the lifetime for the $a\ ^3\Pi\ v = 0$ level. In their study, Eidelsberg et al. (2006) have also reported predissociation rates for B and W states; the large rates found for the $B\ ^1\Sigma^+\ v = 6$ level result from interactions with the D' state. A number of other experimental and theoretical efforts describing the $B - D'$ interaction have been published (Andric et al. 2004; Grozdanov et al. 2004; Baker 2005a; Bitencourt et al. 2007). The triplet k and c states have been the focus of studies (Baker 2005b; Baker & Launay 2005a,b,c) on perturbations caused with other states, while Ben et al. (2007) have described their experiment on perturbations between the a and d levels. Finally, a review (Lefebvre-Brion & Lewis 2007) discussing perturbations in the isoelectronic molecules, CO and N_2 , has also appeared. More spectroscopic work on N_2 is provided below.

New spectroscopic studies on CH, CH^+ , NH have appeared since 2003. Extending the study of Watson (2001) on Rydberg transitions in CH, Sheffer & Federman (2007) have inferred oscillator strengths and predissociation rates for the $3d - X$, $4d - X$, $F - X$, and $D - X$ bands seen in interstellar spectra acquired with the *Hubble Space Telescope*. Theoretical efforts on CH include structure calculations of the $3d$ complex (Vázquez et al. 2007) and oscillator strengths for Rydberg transitions (Lavín et al. 2009); the latter study also determined photoionization cross sections. A new analysis of the $A\ ^1\Pi - X\ ^1\Sigma^+$ system in CH^+ has been completed (Hakalla et al. 2006), and Weselak et al. (2009) have obtained oscillator strengths for the (1,0), (2,0), (3,0), and (4,0) bands of this system from ground-based interstellar spectra. CH^+ photodissociation cross sections have been a focus of recent theoretical work (Barinovs & van Hemert 2004; Bouakline et al. 2005). Revised term values have been determined for the $X\ ^3\Sigma^-$ and $A\ ^3\Pi$ states of NH by Ram & Bernath (2010), who analyzed spectra from the ACE and ATMOS instruments.

Other molecules of interest to interstellar and cometary studies, such as C_2 , CN, and C_3 , have been studied recently. Theoretical efforts have computed oscillator strengths and radiative lifetimes for singlet (Phillips) and triplet (Swan, Ballik-Ramsay, and $d - c$) systems in C_2 (Kokkin et al. 2007; Schmidt & Bacskay 2007); of particular note is that a self-consistent set of oscillator strengths for the Phillips ($A - X$) system of bands is emerging. Experimental studies involving triplet states have been performed with a variety of techniques (Joester et al. 2007; Tanabashi et al. 2007; Nakajima et al. 2009; Bornhauser et al. 2010, 2011), many times focusing on the presence of perturbations. Toffoli & Lucchese (2004) have calculated near threshold photoionization cross sections for C_2 . Spectroscopic studies of the violet and red band systems for isotopologues of CN (Ram et al. 2006, 2010b,c) have improved the precision of the molecular constants for this molecule. This work will also be of interest to stellar astronomy. Shi et al. (2010) have obtained spectroscopic parameters in the isotopologues of CN through *ab initio* calculations. Moreover, experimental work on bands in the $A\ ^1\Pi_u - X\ ^1\Sigma_g^+$ system of C_3 has appeared (McCall et al. 2003; Tanabashi et al. 2005; Zhang et al. 2005; Chen et al. 2010, 2011). Perturbations have been seen in many of these bands. Zhang et al. (2005) have measured lifetimes as well.

Spectroscopic studies on H_2O and HCl have also been conducted. Fillion et al. (2004)

used synchrotron radiation to examine the intense Rydberg nd series in H_2O for energies up to 12 eV. Another study with synchrotron radiation found several new vibrational progressions around 10 eV and also reported absolute cross sections (Mota et al. 2005). Additional experimental efforts obtained absolute absorption cross sections with a synchrotron source (Cheng et al. 2004) and oscillator strengths via electron impact-excitation (Thorn et al. 2007); Cheng et al. (2004) compared their results with additional theoretical computations. Borges (2006b) extended his earlier calculations (Borges 2006a) in a study of oscillator strengths for the $\tilde{A}^1B_1 - \tilde{X}^1A_1$ transition in H_2O . Another electron scattering experiment (Li et al. 2006) derived oscillator strengths for valence-shell excitations in HCl.

4.2. Late-type stars

The transition from metal oxides to metal hydrides is a signature of the latest stellar spectral types. During the reporting period papers on line lists, line strengths, and opacities have appeared. For metal hydrides, these studies include high-resolution spectra acquired with a Fourier transform spectrometer of bands in the $A^2\Pi - X^2\Sigma^+$ and $B^2\Sigma^+ - X^2\Sigma^+$ systems of ^{24}MgH (Shayesteh et al. 2007) and of the $E^2\Pi - X^2\Sigma^+$ transition in CaH and CaD (Ram et al. 2011b). The work on MgH provided the data on the highest lying vibrational level in the ground electronic state. A deperturbation analysis of the MgH spectra was recently completed (Shayesteh & Bernath 2011). Using laser-induced fluorescence, Chowdhury et al. (2006) examined the (1,0) band of the $A^6\Sigma^+ - X^6\Sigma^+$ transition in CrH. New theoretical calculations have led to improved line lists and opacities for transitions in TiH (Burrows et al. 2005) and the $F^4\Delta_i - X^4\Delta_i$ transitions in FeH (Dulick et al. 2003). Magnetic properties of FeH involving the F and X states, including a study of the Zeeman effect, have been determined (Harrison et al. 2008; Harrison & Brown 2008).

4.3. Planetary atmospheres

Nitrogen and sulfur dioxide have received a considerable amount of attention recently. Spectroscopic work on N_2 (Sprengers et al. 2003, 2005b; Lewis et al. 2008a; Vieitez et al. 2008) has focused on extreme UV transitions, where perturbations and predissociation play a significant role; the study by Sprengers et al. (2003) considered the $^{15}\text{N}_2$ and $^{14}\text{N}^{15}\text{N}$ isotopologues. A theoretical calculation (Lewis et al. 2008b) has studied the perturbations involving these high-lying Rydberg states. Oscillator strengths and line widths have been obtained from a number of experimental (Stark et al. 2005, 2008; Heays et al. 2009; Huber et al. 2009) and theoretical (Jungen et al. 2003; Haverd et al. 2005; Lavín & Velasco 2011) efforts. Lifetimes of Rydberg states have been determined through a combination of experiments and calculations (Sprengers et al. 2004; Lewis et al. 2005a,b; Sprengers et al. 2005a; Sprengers & Ubachs 2006). As for SO_2 , the focus has been on absorption cross sections at UV wavelengths (Rufus et al. 2003, 2009; Danielache et al. 2008; Hermans et al. 2009; Blackie et al. 2011), where Danielache et al. (2008) studied isotopologues for sulfur. Furthermore, absorption cross sections for ammonia and its isotopologues (Cheng et al. 2006; Wu et al. 2007) and carbon dioxide (Stark et al. 2007) have been determined experimentally.

Steven R. Federman
chair of Working Group

References

- Allen, M.D., Evenson, K.M., & Brown, J.M. 2004, *J. Mol. Spectrosc.*, 227, 13
- Amano, T., & Hirao, T. 2005, *J. Mol. Spectrosc.*, 233, 7
- Amano, T., Hashimoto, K., & Hirao, T. 2006, *J. Mol. Struct.*, 795, 190
- Amano, T. 2008, *J. Chem. Phys.*, 129, 244305
- Amano, T. 2010a, *ApJ*, 716, L1
- Amano, T. 2010b, *A&A*, 516, L4
- Amano, T. 2010c, *J. Mol. Spectrosc.*, 259, 16
- Andric, L., Bouakline, F., Grozdanov, T.P., & McCarroll, R. 2004, *A&A*, 421, 381
- Araki, M., Furuya, T., & Saito, S. 2001, *J. Mol. Spectrosc.*, 210, 132
- Asvany, O., Ricken, O., Müller, H.S.P., Wiedner, M.C., Giesen, T.F., & Schlemmer, S. 2008, *Phys. Rev. Lett.*, 100, 233004
- Bailleux, S., Walters, A., Grigorova, E., & Margulès, L. 2008, *ApJ*, 679, 920
- Baker, J., 2005a, *Chem. Phys. Lett.*, 408, 312
- Baker, J., 2005b, *J. Mol. Spectrosc.*, 234, 75
- Baker, J., & Launay, F. 2005a, *Chem. Phys. Lett.*, 404, 49
- Baker, J., & Launay, F. 2005b, *Chem. Phys. Lett.*, 415, 296
- Baker, J., & Launay, F. 2005c, *J. Chem. Phys.*, 123, 234302
- Barber, R.J., Tennyson, J., Harris, G.J., & Tolchenov, R.N. 2006, *MNRAS*, 368, 1087
- Barinova, G., & van Hemert, M.C. 2004, *Chem. Phys. Lett.*, 399, 406
- Bauschlicher, C.W., et al. 2010, *ApJS*, 189, 341. See <http://www.astrochem.org/pahdb/>
- Belloche, A., Menten, K.M., Comito, C., Müller, H.S.P., Schilke, P., Ott, J., Thorwirth, S., & Hieret, C. 2008, *A&A*, 482, 179
- Belloche, A., Garrod, R.T., Müller, H.S.P., Menten, K.M., Comito, C., & Schilke, P. 2009, *A&A*, 499, 215
- Ben, J., Li, L., Zheng, L., Chen, Y., & Yang, X. 2007, *Chem. Phys.*, 335, 109
- Bernath, P.F., et al. 2005, *Geophys. Res. Lett.*, 32, L15S01. See <http://www.ace.uwaterloo.ca/>
- Bernath, P.F. 2009, *Int. Rev. Phys. Chem.*, 28, 681
- Bernath, P.F., & Colin, R. 2009, *J. Mol. Spectrosc.*, 257, 20
- Bitencourt, A.C.P., Prudente, F.V., & Vianna, J.D.M. 2007, *J. Phys. B*, 40, 2075
- Bizzocchi, L., Cludi, L., & Degli Esposti, C. 2003, *J. Mol. Spectrosc.*, 218, 53
- Bizzocchi, L., Degli Esposti, C., & Dore, L. 2006, *A&A*, 455, 1161
- Bizzocchi, L., Degli Esposti, C., Dore, L., & Kisiel, Z. 2008, *J. Mol. Spectrosc.*, 251, 138
- Blackie, D., Blackwell-Whitehead, R., Stark, G., Pickering, J.C., Smith, P.L., Rufus, J., & Thorne, A.P. 2011, *JGRE*, 116, 03006
- Borges, I. 2006a, *J. Phys. B*, 39, 641
- Borges, I. 2006b, *Chem. Phys.*, 328, 284
- Bornhauser, P., Knopp, G., Gerber, T., Radi, P.P. 2010, *J. Mol. Spectrosc.*, 262, 69
- Bornhauser, P., Sych, Y., Knopp, G., Gerber, T., & Radi, P.P. 2011, *J. Chem. Phys.*, 134, 044302
- Bouakline, F., Grozdanov, T.P., Andric, L., & McCarroll, R. 2005, *J. Chem. Phys.*, 122, 044108
- Brauer, C.S., Pearson, J.C., Drouin, B.J., & Yu, S. 2009, *ApJS*, 184, 133
- Brown, J.M., Körsgen, H., Beaton, S.P., & Evenson, K.M. 2006, *J. Chem. Phys.*, 124, 234309
- Brown, J.M., & Müller, H.S.P. 2009, *J. Mol. Spectrosc.*, 255, 68
- Brünken, S., Müller, H.S.P., Lewen, F., & Winnewisser, G. 2003, *Phys. Chem. Chem. Phys.*, 5, 1515
- Brünken, S., Müller, H.S.P., Lewen, F., & Giesen, T.F. 2005, *J. Chem. Phys.*, 123, 164315
- Brünken, S., Müller, H.S.P., Thorwirth, S., Lewen, F., & Winnewisser, G. 2006, *J. Mol. Struct.*, 780, 3
- Brünken, S., Müller, H.S.P., Endres, C., Lewen, F., Giesen, T., Drouin, B., Pearson, J.C., Mäder, H. 2007a, *Phys. Chem. Chem. Phys.*, 9, 2103
- Brünken, S., Gottlieb, C.A., Gupta, H., McCarthy, M.C., & Thaddeus, P. 2007b, *A&A*, 464, L33
- Brünken, S., Gottlieb, C.A., McCarthy, M.C., & Thaddeus, P. 2009a, *ApJ*, 697, 880
- Brünken, S., Yu, Z., Gottlieb, C.A., McCarthy, M.C., & Thaddeus, P. 2009b, *ApJ*, 706, 1588
- Burrows, A., Dulick, M., Bauschlicher, C.W., Bernath, P.F., Ram, R.S., Sharp, C.M., & Milsom, J.A. 2005, *ApJ*, 624, 988

- Cacciani, P., & Ubachs, W. 2004, *J. Mol. Spectrosc.*, 225, 62
- Cami, J., et al. 2009, *ApJ*, 690, L122
- Cami, J., van Malderen, R., & Markwick, A.J. 2010a, *ApJS*, 187, 409. See <http://www.spectrafactory.net/>
- Cami, J., Bernard-Salas, J., Peeters, E., & Malek, S.E. 2010b, *Science*, 329, 1180
- Caris, M., Lewen, F., Müller, H.S.P., & Winnewisser, G. 2004, *J. Mol. Struct.*, 695–696, 243
- Caris, M., Giesen, T.F., Duan, C., Müller, H.S.P., Schlemmer, S., & Yamada, K.M.T. 2009, *J. Mol. Spectrosc.*, 253, 99
- Carvajal, M., et al. 2009, *A&A*, 500, 1109
- Carvajal, M., Kleiner, I., & Demaison, J. 2010, *ApJS*, 190, 315
- Cazzoli, G., & Puzzarini, C. 2005, *J. Mol. Spectrosc.*, 233, 280
- Cazzoli, G., Puzzarini, C., Tamassia, F., Borri, S., & Bartalini, S. 2006a, *J. Mol. Spectrosc.*, 235, 265
- Cazzoli, G., Cludi, L., & Puzzarini, C. 2006b, *J. Mol. Struct.*, 780, 260
- Cazzoli, G., & Puzzarini, C. 2006, *J. Mol. Spectrosc.*, 239, 64
- Cazzoli, G., & Puzzarini, C. 2008, *A&A*, 487, 1197
- Cazzoli, G., Dore, L., Puzzarini, C., & Gauss, J. 2010a, *Mol. Phys.*, 108, 2335
- Cazzoli, G., Cludi, L., Puzzarini, C., & Gauss, J. 2010b, *A&A*, 509, A1
- Chakrabarti, K., & Tennyson, J. 2006, *J. Phys. B.*, 39, 1485
- Chen, C.-W., Merer, A.J., Chao, J.-M., & Hsu, Y.-C. 2010, *J. Mol. Spectrosc.*, 263, 56
- Chen, K.-S., Zhang, G., Merer, A.J., Hsu, J.-C., & Chen, W.-J. 2011, *J. Mol. Spectrosc.*, 267, 169
- Cheng, B.-M., Chung, C.-Y., Bahou, M., Lee, Y.-P., Lee, L.C., van Harrevelt, R., & van Hemert, M.C. 2004, *J. Chem. Phys.*, 120, 224
- Cheng, B.-M., et al. 2006, *ApJ*, 647, 1535
- Chowdhury, P.K., Merer, A.J., Rixon, S.J., Bernath, P.F., & Ram, R.S. 2006, *Phys. Chem. Chem. Phys.*, 8, 822
- Christen, D., & Müller, H.S.P. 2003, *Phys. Chem. Chem. Phys.*, 5, 3600
- Colin, R., & Bernath, P.F. 2010, *J. Mol. Spectrosc.*, 263, 120
- Cushing, M.C., et al. 2011, *ApJ*, accepted, arXiv:1108.4678
- Danielache, S.O., Eskebjerg, C., Johnson, M.S., Ueno, Y., & Yoshida, N. 2008, *JGRD*, 113, 17314
- Demyk, K., et al. 2007, *A&A*, 466, 255
- Dickenson, G.D., Nortje, A.C., Steenkamp, C.M., Rohwer, E.G., & du Plessis, A. 2010, *ApJ*, 714, L268
- Dore, L., Bizzocchi, L., Degli Esposti, C., & Tinti, F. 2009, *A&A*, 496, 275
- Dore, L., Bizzocchi, L., Degli Esposti, C., & Gauss, J. 2010, *J. Mol. Spectrosc.*, 263, 44
- Drouin, B.J., Pearson, J.C., Walters, A., & Lattanzi, V. 2006, *J. Mol. Spectrosc.*, 240, 227
- Drouin, B.J., Yu, S., Pearson, J.C., & Müller, H.S.P. 2009, *JQSRT*, 110, 2077
- du Plessis, A., Rohwer, E.G., & Steenkamp, C.M. 2006, *ApJS*, 165, 432
- du Plessis, A., Rohwer, E.G., & Steenkamp, C.M. 2007, *J. Mol. Spectrosc.*, 243, 124
- Duan, Y.-B., Ozier, I., Tsunekawa, S., & Takagi, K. 2003, *J. Mol. Spectrosc.*, 218, 95
- Dubernet, M.L., Boudon, V., Culhane, J.L., et al. 2010, *JQSRT*, 111, 2151
- Dulick, M., Bauschlicher, C.W., Burrows, A., Sharp, C.M., Sharp, R.S., Ram, R.S., & Bernath, P.F. 2003, *ApJ*, 594, 651
- Eidelsberg, M., & Rostas, F. 2003, *ApJS*, 145, 89
- Eidelsberg, M., Launay, F., Ito, K., Matsui, T., Hinnen, P.C., Reinhold, E., Ubachs, W., & Huber, K.P. 2004a, *J. Chem. Phys.*, 121, 292
- Eidelsberg, M., Lemaire, J.L., Fillion, J.H., Rostas, F., Federman, S.R., & Sheffer, Y. 2004b, *A&A*, 424, 355
- Eidelsberg, M., Sheffer, Y., Federman, S.R., Lemaire, J.L., Fillion, J.H., Rostas, F., & Ruiz, J. 2006, *ApJ*, 647, 1543
- Elkeurti, M., Coudert, L.H., Orphal, J., Wlodarczak, G., Fellows, C.E., & Toumi, S. 2008, *J. Mol. Spectrosc.*, 251, 90

12DIVISION XII/COMMISSION 14/WORKING GROUP ON MOLECULAR DATA

- Endres, C.P., Müller, H.S.P., Brünken, S., Paveliev, D.G., Giesen, T.F., Schlemmer, S., & Lewen, F. 2006, *J. Mol. Struct.*, 795, 242
- Endres, C.P., Drouin, B.J., Pearson, J.C., Müller, H.S.P., Lewen, F., Schlemmer, S., & Giesen, T. F. 2009, *A&A*, 504, 635
- Fillion, J.-H., Ruiz, J., Yang, X.-F., Castillejo, M., Rostas, F., & Lemaire, J.-L. 2004, *J. Chem. Phys.*, 120, 6531
- Fisher, J., Paciga, G., Xu, L.-H., Zhao, S.B., Moruzzi, G., & Lees, R.M. 2007, *J. Mol. Spectrosc.*, 245, 7
- Flores-Mijangos, J., Brown, J.M., Matsushima, F., Odashima, H., Takagi, K., Zink, L.R., & Evenson, K.M. 2004, *J. Mol. Spectrosc.*, 225, 189
- Flory, M.A., Halfen, D.T., & Ziurys, L.M. 2004, *J. Chem. Phys.*, 121, 8385
- Flory, M.A., McLamarrah, S.K., & Ziurys, L.M. 2006, *J. Chem. Phys.*, 125, 194304
- Flory, M.A., & Ziurys, L.M. 2008, *J. Mol. Spectrosc.*, 247, 76
- Fortman, S.M., Medvedev, I.R., Neese, C.F., & De Lucia, F.C. 2010, *ApJ*, 725, 1682
- Fortman, S.M., Medvedev, I.R., Neese, C.F., & De Lucia, F.C. 2011, *ApJ*, 737, 20
- Fujimori, R., Kawaguchi, K., & Amano, T. 2011, *ApJ*, 729, L2
- Galué, H.A., Rice, C.A., Steill, J.D., Oomens, J. 2011, *J. Chem. Phys.*, 134, 054310
- Gendriesch, R., Lewen, F., Klapper, G., Menten, K.M., Winnewisser, G., Coxon, J.A., & Müller, H.S.P. 2009, *A&A*, 497, 927
- Gilijamse, J.J., Hoekstra, S., Meek, S.A., Metsälä, M., van de Meerakker, S.Y.T., Meijer, G., & Groenenboom, G.C. 2007, *J. Chem. Phys.*, 127, 221102
- Golubiatnikov, G.Y., Markov, V.N., Guarnieri, A., & Knöchel, R. 2006, *J. Mol. Spectrosc.*, 240, 251
- Gottlieb, C.A., Brünken, S., McCarthy, M.C., & Thaddeus, P. 2007, *J. Chem. Phys.*, 126, 191101
- Gottlieb, C.A., McCarthy, M.C., & Thaddeus, P. 2010, *ApJS*, 189, 261
- Groner, P., Medvedev, I.R., De Lucia, F.C., & Drouin, B.J. 2008, *J. Mol. Spectrosc.*, 251, 180
- Grozdanov, T.P., Bouakline, F., Andric, L., & McCarroll, R. 2004, *J. Phys. B.*, 37, 1737
- Gupta, H., Brünken, S., Tamassia, F., Gottlieb, C.A., McCarthy, M.C., & Thaddeus, P. 2007, *ApJ*, 655, L57
- Hakalla, R., Kepa, R., Szajna, W., & Zachwieja, M. 2006, *Eur. Phys. J. D*, 38, 481
- Halfen, D.T., & Ziurys, L.M. 2004, *ApJ*, 607, L63
- Halfen, D.T., & Ziurys, L.M. 2008, *ApJ*, 672, L77
- Halfen, D.T., Ziurys, L.M., Pearson, J.C., & Drouin, B.J. 2008, *ApJ*, 687, 731
- Halfen, D.T., Sun, M., Clouthier, D.J., & Ziurys, L.M. 2009a, *J. Chem. Phys.*, 130, 014305
- Halfen, D.T., Ziurys, L.M., & Brown, J.M. 2009b, *J. Chem. Phys.*, 130, 164301
- Hargreaves, R.J., Li, G., & Bernath, P.F. 2011, *ApJ*, 735, 111
- Harrison, J.J., & Brown, J.M. 2008, *ApJ*, 686, 1426
- Harrison, J.J., Brown, J.M., Halfen, D.T., & Ziurys, L.M. 2006, *ApJ*, 637, 1143
- Harrison, J.J., Brown, J.M., Flory, M.A., Sheridan, P.M., McLamarrah, S.K., & Ziurys, L.M. 2007, *J. Chem. Phys.*, 127, 194308
- Harrison, J.J., Brown, J.M., Chen, J., Steimle, T.C., & Sears, T.J. 2008, *ApJ*, 679, 854
- Hase, F., Wallace, L., McLeod, S.D., Harrison, J.J., & Bernath, P.F. 2010, *JQSRT*, 111, 521.
See <http://www.ace.uwaterloo.ca/solaratlas.html/>
- Haverd, V.E., Lewis, B.R., Gibson, S.T., & Stark, G. 2005, *J. Chem. Phys.*, 123, 214304
- Heays, A.N., Lewis, B.R., Stark, G., Yoshino, K., Smith, P.L., Huber, K.P., & Ito, K. 2009, *J. Chem. Phys.*, 131, 194308
- Hermans, C., Vandaele, A.C., & Fally, S. 2009, *JQSRT*, 110, 756
- Hirata, Y., Kubota, S., Watanabe, S., Momose, T., & Kawaguchi, K. 2008, *J. Mol. Spectrosc.*, 251, 314
- Huang, X., Schwenke, D.W., & Lee, T.J. 2011, *J. Chem. Phys.*, 134, 044320 and 044321
- Huber, K.P., Chan, M.-C., Stark, G., Ito, K., & Matsui, T. 2009, *J. Chem. Phys.*, 131, 084301
- Hübers, H.-W., Evenson, K.M., Hill, C., & Brown, J.M. 2009, *J. Chem. Phys.*, 131, 034311
- Hudgins, D.M., Bauschlicher, C.W., Allamandola, L.J. 2005, *ApJ*, 632, 316
- Iglesias-Groth, S., Cataldo, F., & Manchado, A. 2011, *MNRAS*, 413, 213

- Ilyushin, V.V., Alekseev, E.A., Dyubko, S.F., Kleiner, I., & Hougen, J.T. 2004, *J. Mol. Spectrosc.*, 227, 115
- Ilyushin, V.V., Alekseev, E.A., Dyubko, S.F., Motiyenko, R.A., & Lovas, F.J. 2005, *J. Mol. Spectrosc.*, 231, 15
- Ilyushin, V., & Lovas, F.J. 2007, *J. Phys. Chem. Ref. Data*, 36, 1141
- Ilyushin, V., Kleiner, I., & Lovas, F.J. 2008, *J. Phys. Chem. Ref. Data*, 37, 97
- Ilyushin, V., Kryvda, A., & Alekseev, E. 2009, *J. Mol. Spectrosc.*, 255, 32
- Jacquinet-Husson, N., Crepeau, L., Armante, R., et al. 2011, *JQSRT*, 112, 2395
- Jansen, P., Xu, L.-H., Kleiner, I., Ubachs, W., & Bethlem, H.L. 2011, *Phys. Rev. Lett.*, 106, 100801
- Joester, J.A., Nakajima, M., Reilly, N.J., Kokkin, D.L., Nauta, K., Kable, S.H., & Schmidt, T.W. 2007, *J. Chem. Phys.*, 127, 214303
- Jungen, Ch., Huber, K.P., Jungen, M., & Stark, G. 2003, *J. Chem. Phys.*, 118, 4517
- Kania, P., Hermanns, M., Brünken, S., Müller, H.S.P., & Giesen, T.F. 2011, *J. Mol. Spectrosc.*, 268, 173
- Kassi, S., & Campargue, A. 2011, *J. Mol. Spectrosc.*, 267, 36
- Kawaguchi, K., Kasai, Y., Ishikawa, S.-I., & Kaifu, N. 1995, *PASJ*, 47, 853
- Kawahara, H., Kato, H., Hoshino, M., Tanaka, H., & Brunger, M.J. 2008, *Phys. Rev.*, A77, 012713
- Killian, T.C., Gottlieb, C.A., & Thaddeus, P. 2007, *J. Chem. Phys.*, 127, 114320
- Kisiel, Z., Pszczółkowski, L., Drouin, B.J., Brauer, C.S., Yu, S., & Pearson, J.C. 2009, *J. Mol. Spectrosc.*, 258, 26
- Kisiel, Z., et al. 2010, *Phys. Chem. Chem. Phys.*, 12, 8329
- Kokkin, D.L., Bacskay, G.B., & Schmidt, T.W. 2007, *J. Chem. Phys.*, 126, 084302
- Kryvda, A.V., Gerasimov, V.G., Dyubko, S.F., Alekseev, E.A., & Motiyenko, R.A. 2009, *J. Mol. Spectrosc.*, 254, 28
- Lattanzi, V., Walters, A., Drouin, B.J., & Pearson, J.C. 2008, *ApJS*, 176, 536
- Lattanzi, V., Gottlieb, C.A., Thaddeus, P., Thorwirth, S., & McCarthy, M.C. 2010, *ApJ*, 720, 1717
- Lauvergnat, D., Coudert, L.H., Klee, S., & Smirnov, M. 2009, *J. Mol. Spectrosc.*, 256, 204
- Lavín, C., Velasco, A.M., Martín, I. 2009, *ApJ*, 692, 1354
- Lavín, C., & Velasco, A.M. 2011, *ApJ*, 739, 16
- Lefebvre-Brion, H., & Lewis, B.R., 2007, *Mol. Phys.*, 105, 1625
- Lefebvre-Brion, H., Liebermann, H.P., & Vázquez, G.J. 2010, *J. Chem. Phys.*, 132, 024311
- Levshakov, S.A., Kozlov, M.G., & Reimers, D. 2011, *ApJ*, 738, 26
- Lewis, B.R., Gibson, S.T., Zhang, W., Lefebvre-Brion, H., & Robbe, J.-M. 2005a, *J. Chem. Phys.*, 122, 144302
- Lewis, B.R., Gibson, S.T., Sprengers, J.P., Ubachs, W., Johansson, A., & Wahlström, C.-G. 2005b, *J. Chem. Phys.*, 123, 236101
- Lewis, B.R., Baldwin, K.G.H., Sprengers, J.P., Ubachs, W., Stark, G., & Yoshino, K. 2008a, *J. Chem. Phys.*, 129, 164305
- Lewis, B.R., Heays, A.N., Gibson, S.T., Lefebvre-Brion, H., & Lefebvre, R. 2008b, *J. Chem. Phys.*, 129, 164306
- Li, G., Gordon, I.E., Bernath, P.F., & Rothman, L.S. 2011, *JQSRT*, 112, 1543
- Li, W.-B., Zhu, L.-F., Yuan, Z.-S., Liu, X.-J., & Xu, K.-Z. 2006, *J. Chem. Phys.*, 125, 154310
- Livingston, W., & Wallace, L. 2003, *An Atlas of the Solar Spectrum in the Infrared from 1850 – 9000 cm⁻¹ (1.1 to 5.4 μm)*, Revised, N.S.O. Technical Report # 03-001, National Solar Observatory, Tucson. See <ftp://nsokp.nso.edu/pub/atlas/photat/>
- Lovas, F.J. 2004, *J. Phys. Chem. Ref. Data*, 33, 177
- Lovas, F.J., & Groner, P. 2006, *J. Mol. Spectrosc.*, 236, 173
- Lovas, F.J., Plusquellic, D.F., Pate, B.H., Neill, J.L., Muckle, M.T., & Remijan, A.J. 2009, *J. Mol. Spectrosc.*, 257, 82
- Maeda, A., Hirao, T., Bernath, P.F., & Amano, T. 2001, *J. Mol. Spectrosc.*, 210, 250
- Maeda, A., et al. 2008, *ApJS*, 176, 543
- Margulès, L., Lewen, F., Winnewisser, G., Botschwina, P., & Müller, H.S.P. 2003, *Phys. Chem. Chem. Phys.*, 5, 2770

14DIVISION XII/COMMISSION 14/WORKING GROUP ON MOLECULAR DATA

- Margulès, L., Coudert, L.H., Møllendal, H., Guillemin, J.-C., Huet, T.R., & Janečková, R. 2009a, *J. Mol. Spectrosc.*, 254, 55
- Margulès, L., et al. 2009b, *A&A*, 493, 565
- Margulès, L., et al. 2010, *ApJ*, 714, 1120
- McCall, B.J., Casaes, R.N., Ádámkóvics, M., & Saykally, R.J. 2003, *Chem. Phys. Lett.*, 374, 583
- McCarthy, M.C., & Thaddeus, P. 2005, *J. Chem. Phys.*, 122, 174308
- McCarthy, M.C., Gottlieb, C.A., Gupta, H., & Thaddeus, P. 2006, *ApJ*, 652, L141
- McLamarrah, S.K., Sheridan, P.M., & Ziurys, L.M. 2005, *Chem. Phys. Lett.*, 414, 301
- Medvedev, I.R., De Lucia, F.C., & Herbst, E. 2009, *ApJS*, 181, 433
- Mladenović, M., Lewerenz, M., McCarthy, M.C., & Thaddeus, P. 2009, *J. Chem. Phys.*, 131, 174308
- Mota, R., et al. 2005, *Chem. Phys. Lett.*, 416, 152
- Mukhopadhyay, I., Perry, D.S., Duan, Y.-B., Pearson, J.C., Albert, S., Butler, R.A.H., Herbst, E., & De Lucia, F.C. 2002, *J. Chem. Phys.*, 116, 3710
- Müller, H.S.P., & Christen, D. 2004, *J. Mol. Spectrosc.*, 228, 298
- Müller, H.S.P., Menten, K.M., & Mäder, H. 2004, *A&A*, 428, 1019
- Müller, H.S.P., & Brünken, S. 2005, *J. Mol. Spectrosc.*, 232, 213
- Müller, H.S.P., Schlöder, F., Stutzki, J., & Winnewisser, G. 2005, *J. Mol. Struct.*, 742, 215
- Müller, H. S. P., et al. 2007, *Phys. Chem. Chem. Phys.*, 9, 1579
- Müller, H.S.P., Drouin, B.J., & Pearson, J. C. 2009, *A&A*, 506, 1487
- Müller, H.S.P., Dong, F., Nesbitt, D.J., Furuya, T., & Saito, S. 2010, *Phys. Chem. Chem. Phys.*, 12, 8362
- Müller, H.S.P. 2010, *A&A*, 514, L6
- Müller, H.S.P., Coutens, A., Walters, A., Grabow, J.-U., & Schlemmer, S. 2011, *J. Mol. Spectrosc.*, 267, 100
- Nakajima, M., Joester, J.A., Page, N.I., Reilly, N.J., Bacskay, G.B., Schmidt, T.W., & Kable, S.H. 2009, *J. Chem. Phys.*, 131, 044301
- Nassar, R., & Bernath, P.F. 2003, *JQSRT*, 82, 279
- Nemes, L. et al. 1994, *Chem. Phys. Lett.*, 218, 295
- Ozeki, H., Bailleux, S., & Włodarczak, G. 2011, *A&A*, 527, A64
- Pearson, J.C., & Drouin, B.J. 2006, *ApJ*, 647, L83
- Pearson, J.C., Brauer, C.S., & Drouin, B.J. 2008, *J. Mol. Spectrosc.*, 251, 394
- Pearson, J.C., Brauer, C.S., Drouin, B.J., & Xu, L.-H. 2009, *Can. J. Phys.*, 87, 449
- Plusquellic, D.F., Lovas, F.J., Pate, B.H., Neill, J.L., Muckle, M.T., & Remijan, A.J. 2009, *J. Phys. Chem. A*, 1131, 12911
- Polehampton, E.T., Brown, J.M., Swinyard, B.M., & Baluteau, J.-P. 2003, *A&A*, 406, L47
- Puzzarini, C., Dore, L., & Cazzoli, G. 2003, *J. Mol. Spectrosc.*, 217, 19
- Puzzarini, C., Cazzoli, G., Harding, M.E., Vázquez, J., & Gauss, J. 2009, *J. Chem. Phys.*, 131, 234304
- Ram, R.S., Davis, S.P., Wallace, L., Engleman, R., Appadoo, D.R.T., & Bernath, P.F. 2006, *J. Mol. Spectrosc.*, 237, 225
- Ram, R.S., & Bernath, P.F. 2010, *J. Mol. Spectrosc.*, 260, 115
- Ram, R.S., Wallace, L., & Bernath, P.F. 2010b, *J. Mol. Spectrosc.*, 263, 82
- Ram, R.S., Wallace, L., Hinkle, K., & Bernath, P.F. 2010c, *ApJS*, 188, 500
- Ram, R.S., & Bernath, P.F. 2011a, *ApJS*, 194, 34
- Ram, R.S., Tereszchuk, K., Gordon, I.E., Walker, K.A., & Bernath, P.F. 2011b, *J. Mol. Spectrosc.*, 266, 86
- Ricks, A.M., Douberly, G.E., Duncan, M.A. 2009, *ApJ*, 702, 301
- Rothman, L.S., et al. 2009, *JQSRT*, 110, 523. See <http://www.cfa.harvard.edu/HITRAN/>
- Rothman, L.S., et al. 2010, *JQSRT*, 111, 2139
- Rufus, J., Stark, G., Smith, P.L., Pickering, J.C., & Thorne, A.P. 2003, *JGRE*, 108, 5011
- Rufus, J., Stark, G., Thorne, A.P., Pickering, J.C., Blackwell-Whitehead, R.J., Blackie, D., & Smith, P.L. 2009, *JGRE*, 114, 06003
- Savin, D.W., et al. 2011, *Prog. Phys.*, in press
- Schmidt, T.W., & Bacskay, G.B. 2007, *J. Chem. Phys.*, 127, 234310

- Sharpe, S.W., Johnson, T.J., Sams, R.L., Chu, P.M., Rhoderick, G.C., & Johnson, P.A. 2004, *Appl. Spectrosc.*, 58, 1452. See <http://nwir.pnl.gov/>
- Shayesteh, A., & Bernath, P.F. 2011, *J. Chem. Phys.*, 135, 094308
- Shayesteh, A., Tereszchuk, K., Bernath, P.F., & Colin, R. 2003a, *J. Chem. Phys.*, 118, 1158
- Shayesteh, A., Tereszchuk, K., Bernath, P.F., & Colin, R. 2003b, *J. Chem. Phys.*, 118, 3622
- Shayesteh, A., Appadoo, D.R.T., Gordon, I., & Bernath, P.F. 2003c, *J. Chem. Phys.*, 119, 7785
- Shayesteh, A., Appadoo, D.R.T., Gordon, I., LeRoy, R.J., & Bernath, P.F. 2004a, *J. Chem. Phys.*, 120, 10002
- Shayesteh, A., Walker, K.A., Gordon, I., Appadoo, D.R.T., & Bernath, P.F. 2004b, *J. Mol. Struct.*, 695-696, 23
- Shayesteh, A., Henderson, R.D.E., Le Roy, R.J., & Bernath, P.F. 2007, *J. Phys. Chem. A*, 111, 12495
- Sheffer, Y., Federman, S.R., & Andersson, B.-G. 2003, *ApJ*, 597, L29
- Sheffer, Y., & Federman, S.R. 2007, *ApJ*, 659, 1352
- Sheridan, P.M., & Ziurys, L.M. 2003, *J. Chem. Phys.*, 118, 6370
- Sheridan, P.M., McLamarrah, S.K., & Ziurys, L.M. 2003, *J. Chem. Phys.*, 119, 9496
- Sheridan, P.M., Flory, M.A., & Ziurys, L.M. 2004, *J. Chem. Phys.*, 121, 8360
- Shi, D., Liu, H., Zhang, X., Sun, J., Zhu, Z., & Liu, Y. 2010, *J. Mol. Struct.*, 956, 10
- Sprengers, J.P., Ubachs, W., Baldwin, K.G.H., Lewis, B.R., & Tchang-Brillet, W-ÜL. 2003, *J. Chem. Phys.*, 119, 3160
- Sprengers, J.P., Johansson, A., L'Huillier, A., Wahlström, C.-G., Lewis, B.R., & Ubachs, W. 2004, *Chem. Phys. Lett.*, 389, 348
- Sprengers, J.P., Ubachs, W., & Baldwin, K.G.H. 2005a, *J. Chem. Phys.*, 122, 144301
- Sprengers, J.P., Reinhold, E., Ubachs, W., Baldwin, K.G.H., & Lewis, B.R. 2005b, *J. Chem. Phys.*, 123, 144315
- Sprengers, J.P., & Ubachs, W. 2009, *J. Mol. Spectrosc.*, 235, 176
- Stark, G., Huber, K.P., Yoshino, K., Smith, P.L., & Ito, K. 2005, *J. Chem. Phys.*, 123, 214303
- Stark, G., Yoshino, K., Smith, P.L., & Ito, K. 2007, *JQSRT*, 103, 67
- Stark, G., Lewis, B.R., Heays, A.N., Yoshino, K., Smith, P.L., & Ito, K. 2008, *J. Chem. Phys.*, 128, 114302
- Steinmann, C.M., Rohwer, E.G., & Stafast, H. 2003, *ApJ*, 590, L123; errat. *ApJ*, 591, L167
- Sumiyoshi, Y., Ueno, T., & Endo, Y. 2003, *J. Chem. Phys.*, 119, 1426
- Tanabashi, A., Hirao, T., Amano, T., & Bernath, P.F. 2005, *ApJ*, 624, 1116
- Tanabashi, A., Hirao, T., Amano, T., & Bernath, P.F. 2007, *ApJS*, 169, 472
- Tashkun, S.A., & Perevalov, V.I. 2011, *JQSRT*, 112, 1403
- Tenenbaum, E.D., Flory, M.A., Pulliam, R.L., & Ziurys, L.M. 2007, *J. Mol. Spectrosc.*, 244, 153
- Thaddeus, P., Gottlieb, C.A., Gupta, H., Brünken, S., McCarthy, M.C., Agúndez, M., Guélin, M., & Cernicharo, J. 2008, *ApJ*, 677, 1132
- Thorn, P.A., et al. 2007, *J. Chem. Phys.*, 126, 064306
- Tinti, F., Bizzocchi, L., Degli Esposti, C., & Dore, L. 2007, *ApJ*, 669, L113
- Toffoli, D., & Lucchese, R.R. 2004, *J. Chem. Phys.*, 120, 6010
- Vázquez, G.J., Amero, J.M., Liebermann, H.P., Buenker, R.J., Lefebvre-Brion, H. 2007, *J. Chem. Phys.*, 126, 164302
- Vázquez, G.J., Amero, J.M., Liebermann, H.P., & Lefebvre-Brion, H. 2009, *J. Phys. Chem. A*, 113, 13395
- Vieitez, M.O., Ivanov, T.I., de Lange, C.A., Ubachs, W., Heays, A.N., Lewis, B.R., & Stark, G. 2008, *J. Chem. Phys.*, 128, 134313
- Wallace, L., Hinkle, K., & Livingston, W.C. 2001, *Sunspot Umbral Spectra in the Region 4000 to 8640 cm⁻¹ (1.16 to 2.50 μm)*, N.S.O. Technical Report # 01-001, National Solar Observatory, Tucson. See <ftp://nsokp.nso.edu/pub/atlas/spot5atl/>
- Wallace, L., Hinkle, K., & Livingston, W.C. 2002, *Sunspot Umbral Spectra in the Regions 1925 to 2226 and 2392 to 3480 cm⁻¹ (2.87 to 4.18 and 4.48 to 5.35 μm)*, N.S.O. Technical Report # 02-001, National Solar Observatory, Tucson. See <ftp://nsokp.nso.edu/pub/atlas/spot6atl/>
- Walters, A., Müller, H.S.P., Lewen, F., & Schlemmer, S. 2009, *J. Mol. Spectrosc.*, 257, 24
- Watson, J.K.G. 2001, *ApJ*, 555, 472

16 DIVISION XII/COMMISSION 14/WORKING GROUP ON MOLECULAR DATA

- Weselak, T., Galazutdinov, G.A., Musaev, F.A., Beletsky, Y., & Krelowski, J. 2009, *A&A*, 495, 189
- Wu, Y.-J., Lu, H.-C., Chen, H.-K., Cheng, B.-M., Lee, Y.-P., & Lee, L.C. 2007, *J. Chem. Phys.*, 127, 154311
- Xu, L.-H., et al. 2008, *J. Mol. Spectrosc.*, 251, 305
- Yang, X., Ben, J., Li, L., & Chen, Y. 2008, *JQSRT*, 109, 468
- Yonezu, T., Matsushima, F., Moriwaki, Y., Takagi, K., & Amano, T. 2009, *J. Mol. Spectrosc.*, 256, 238
- Yoshikawa, T., Sumiyoshi, Y., & Endo, Y. 2009a, *J. Chem. Phys.*, 130, 094302
- Yoshikawa, T., Sumiyoshi, Y., & Endo, Y. 2009b, *J. Chem. Phys.*, 130, 164303
- Yu, S., Drouin, B.J., Pearson, J.C., & Pickett, H.M. 2009, *ApJS*, 180, 119
- Yu, S., et al. 2010, *J. Chem. Phys.*, 133, 174317
- Yurchenko, S.N., Barber, R.J., & Tennyson, J. 2011, *MNRAS* 413, 1828
- Zack, L.N., Pulliam, R.L., & Ziurys, L.M. 2009, *J. Mol. Spectrosc.*, 256, 186
- Zelinger, Z., Amano, T., Ahrens, V., Brünken, S., Lewen, F., Müller, H.S.P., & Winnewisser, G. 2003, *J. Mol. Spectrosc.*, 220, 223
- Zhang, G., Chen, K.-S., Merer, A.J., Hsu, Y.-C., Chen, W.-J., Shaji, S., & Liao, Y.-A. 2005, *J. Chem. Phys.*, 122, 244308
- Zobov, N.F., et al. 2008, *MNRAS*, 387, 1093